This is a “Bring Back Quiz”. It is due March 24th. Show all work to receive credit.

1. (3 Pts) What is the resulting concentration when 455.8 mL of a 0.0786 M Na₂SO₄ solution is evaporated to a volume of 50.00 mL?

2. (3 Pts) What concentration H₃PO₄ results when 50.00 mL of 0.355 M H₃PO₄ solution is diluted to 400.0 mL?

3. (4 Pts) How many grams of HNO₃ are present in 450.0 mL of 0.0550 M HNO₃ solution?

4. 25.00 mL of 0.505 M hydrochloric acid solution is reacted with 20.50 mL of 0.303 M barium hydroxide solution.
   a. (5 Pts) Determine how many moles of the excess reactant is present when the reaction is done.
   b. (5 Pts) Determine the concentration (in moles per liter) of the remaining (excess) reactant.

5. (5 Pts) A barium hydroxide solution is being standardized with potassium hydrogen phthalate (KHP). If it took 33.25 mL of the barium hydroxide solution to neutralize 0.5728 grams KHP, what was the molarity of the barium hydroxide solution?